Electrochemistry Quiz **Name:** _____

May I post your solution? [ ] Yes [ ] No [ ] Yes, but redact my name

Consider a Galvanic/voltaic cell based on the following half reactions under standard conditions.

$$\mathbf{Fe^{2+}\,(aq) + 2\,e^- \rightarrow Fe\,(s)} \qquad \boldsymbol{E^\circ = -0.440\ V}$$

$$\mathbf{Cd^{2+}\,(aq) + 2\,e^- \rightarrow Cd\,(s)} \qquad \boldsymbol{E^\circ = -0.403\ V}$$

What will be the cell potential (E_{cell}) when the cathode solution *changes* by 0.827 M?
You may leave your answer in the form of an expression.

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