

## Equilibrium Quiz

Name: \_\_\_\_\_

May I post your solution?

Yes

No

Yes, but redact my name

Consider the reaction:



At 35 °C, the equilibrium constant is  $K_c = 1.6 \times 10^{-5}$ . In an experiment, you place 1.0 mol of NO (g) and 1.0 mol of Cl<sub>2</sub> (g) into a 2.0 L container and allow the system to reach equilibrium.

Set up an ICE chart and an expression that would allow you to calculate the equilibrium concentration of NO (g).

---

## Equilibrium Quiz

Name: \_\_\_\_\_

May I post your solution?

Yes

No

Yes, but redact my name

Consider the reaction:



At 35 °C, the equilibrium constant is  $K_c = 1.6 \times 10^{-5}$ . In an experiment, you place 1.0 mol of NO (g) and 1.0 mol of Cl<sub>2</sub> (g) into a 2.0 L container and allow the system to reach equilibrium.

Set up an ICE chart and an expression that would allow you to calculate the equilibrium concentration of NO (g).