



EXPERIMENT 2

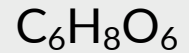
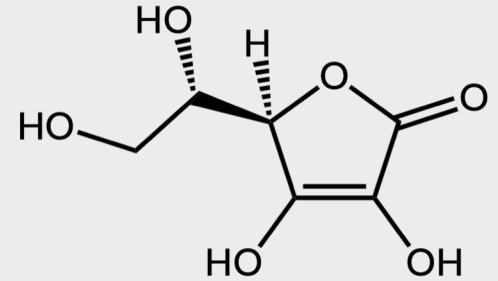
QUANTITATIVE DETERMINATION OF
VITAMIN C IN COMMERCIAL TABLETS

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Purpose

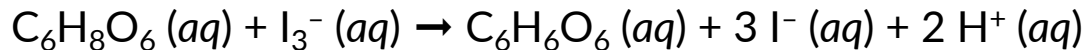
To determine the mass of vitamin C (ascorbic acid) in a commercial tablet.



Method

Titrimetry (volumetric analysis), which requires a 'clean' chemical reaction that involves the analyte as one of the reactants and has a well-defined stoichiometry. We also need a good titrant and indicator.

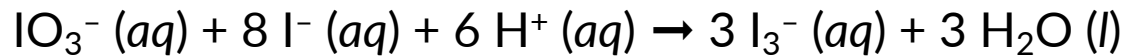
Ascorbic acid is a reductant. As such, it undergoes oxidation in the presence of triiodide (I_3^-) via:



oxidation

Problem → No triiodide salts are commercially available.

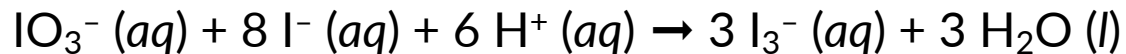
Solution → We can make I_3^- readily in the lab.



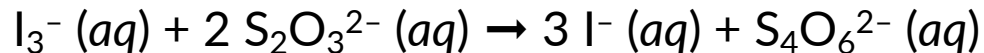
Procedural Outline

The main idea in the titrimetric determination of ascorbic acid is to:

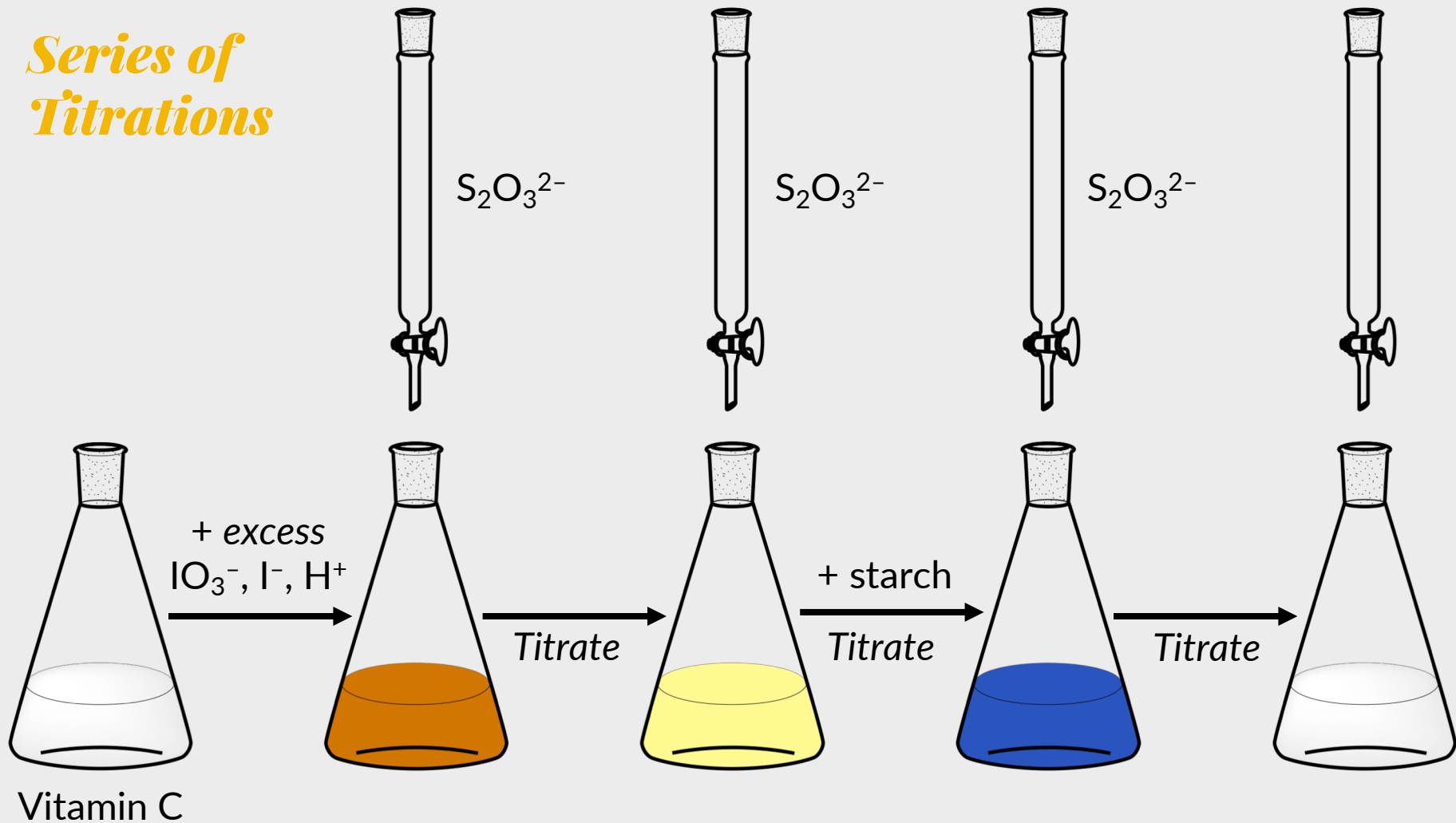
Make an excess of I_3^- using IO_3^- in the presence of ascorbic acid.



Titrate the *remaining* I_3^- with thiosulfate ($S_2O_3^{2-}$) using starch indicator.



Series of Titrations



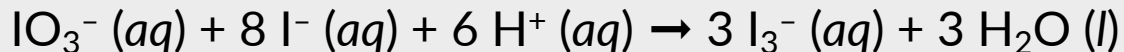
Standardize the thiosulfate solution

Primary standards: very stable, high purity, and have known molar mass

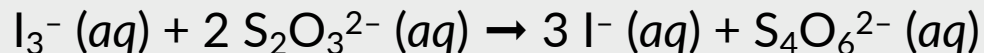
Commercially available hydrates of $\text{Na}_2\text{S}_2\text{O}_3$ are not primary standards because they may have different hydrates ($\text{Na}_2\text{S}_2\text{O}_3 \cdot x\text{H}_2\text{O}$).



KIO_3 is a primary standard. It can be used to standardize thiosulfate solutions.



Titrate the I_3^- with the *same* thiosulfate ($\text{S}_2\text{O}_3^{2-}$) solution.



Now we can determine the concentration of thiosulfate solution used.

Notes

1. Work independently today.
2. Try to finish standardization by 3:45pm today.
3. There are worked-out examples in the lab manual.
4. Lab report: Ex. 3 due next week Tuesday/Wednesday
5. Quizzes: use a pen + wait