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Name: Maya Sanghvi

yes, not 6/22/18

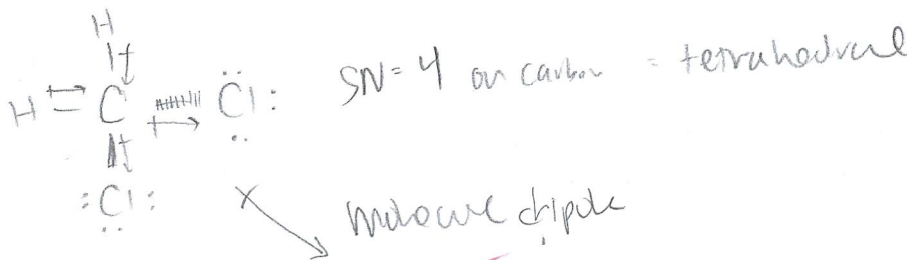
Periodic Table of Elements

1 H 1.008																	2 He 4.003
3 Li 6.941	4 Be 9.012											5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 19.00	10 Ne 20.18
11 Na 22.99	12 Mg 24.31											13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.06	17 Cl 35.45	18 Ar 39.95
19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.88	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.69	29 Cu 63.55	30 Zn 65.38	31 Ga 69.72	32 Ge 72.59	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80
37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc (98)	44 Ru 101.1	45 Rh 102.9	46 Pd 106.4	47 Ag 107.9	48 Cd 112.4	49 In 114.8	50 Sn 118.7	51 Sb 121.8	52 Te 127.6	53 I 126.9	54 Xe 131.3
55 Cs 132.9	56 Ba 137.3	57 La 138.9	72 Hf 178.5	73 Ta 180.9	74 W 183.9	75 Re 186.2	76 Os 190.2	77 Ir 192.2	78 Pt 195.1	79 Au 197.0	80 Hg 200.6	81 Tl 204.4	82 Pb 207.2	83 Bi 209.0	84 Po (209)	85 At (210)	86 Rn (222)
87 Fr (223)	88 Ra 226	89 Ac (227)															

Consider the molecule dichloromethane (CH_2Cl_2); by considering its Lewis structure and geometry you should be able to answer the following questions.

1. Is this molecule polar?
2. What is the hybridization of the orbitals on the C?

You must support your answer to 1 and 2 with a picture of the geometry.



1. Yes, this molecule is polar ✓
2. The hybridization on C is sp^3 because the steric number is 4 ✓

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3Name: Ben Lauring

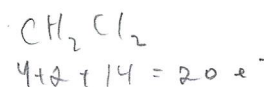
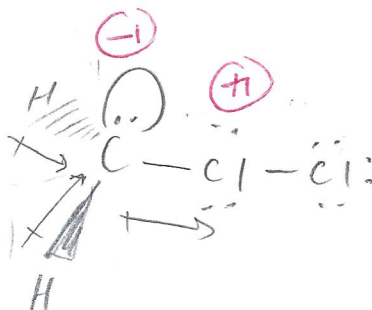
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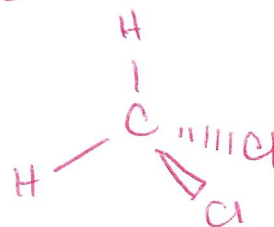
Consider the molecule dichloromethane (CH_2Cl_2); by considering its Lewis structure and geometry you should be able to answer the following questions.

1. Is this molecule polar? *yes* ✓
2. What is the hybridization of the orbitals on the C? *sp^3* ✓

You must support your answer to 1 and 2 with a picture of the geometry.



Not the proper structure of dichloromethane, but okay.



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3Name: Fiona Reed

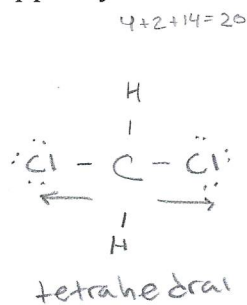
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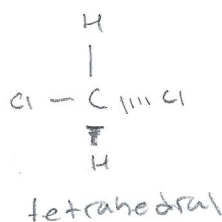
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1. Is this molecule polar?
2. What is the hybridization of the orbitals on the C?

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1. The molecule is polar because ✓
the chlorine molecules are very electronegative and due to the tetrahedral shape their charges will not cancel each other out.



2. C has a steric number of 4 and is tetrahedral so the hybridization is sp^3 . ✓